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Section 8.1 Assessment page 247 7. Identify the type of atom that generally forms covalent bonds. The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule applies to covalent bonds. Atoms share valence

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electrons; the shared electrons complete the octet of each atom.
9. Illustrate the formation of single ...

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Name: _____ Class: _____ Date: _____ Assessment Chemical Bonding Section Quiz: Covalent Bonding and Molecular Compounds In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____

1. If two covalently bonded atoms move closer than a distance of the bond length, the potential energy of the atoms a. becomes negative.

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247 7. Identify the type of atom that generally forms covalent
bonds. The majority of covalent bonds form between nonmetallic
elements. 8. Describe how the octet Page 5/15

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Chemical Bonding SECTION 2 SHORT ANSWER Answer the
following questions in the space provided. 1. Use the concept of
potential energy to describe how a covalent bond forms between
two atoms. As the atoms involved in the formation of a covalent
bond approach each other, the electron-proton attraction is
stronger than the electron-electron and ...

6 Chemical Bonding

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Learn vocabulary, terms, and more with flashcards, games, and other Page 8/29 Chapter 8 Covalent Bonding Study Answers A covalent bond in which three pairs of electrons are shared by two atoms. Chemistry Chapter 8 Flashcards ...

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Section 6.2 – Covalent Bonding. A covalent bond is a chemical bond in which two atoms share a pair of valence electrons. When two atoms share one pair of electrons, the bond is called a single bond.

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